

2. PRECIPITATION RXNS

UNIT 1 AQUEOUS CHEMICAL REACTIONS

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1

REVIEW

Write the dissociation equations for the following solutions and draw a diagram to model the process:

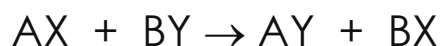
1. Lead(II) nitrate is dissolved in water.
2. Potassium chloride is dissolved in water.

WHAT WOULD HAPPEN IF YOU MIXED THESE TWO SOLUTIONS TOGETHER?

2

DOUBLE REPLACEMENT REACTIONS FORM PRECIPITATES

When two soluble ionic compounds are mixed together in solution, the ions of the compounds **exchange places** to form two new compounds.



One of the compounds formed is usually either:

1. a **precipitate** (an insoluble solid), or
2. an **insoluble gas** that bubbles out of solution.

3

OBSERVE A PRECIPITATION REACTION...

1. Place your Petri dish on a white piece of paper and fill it about half full of distilled water.
2. Have one person place a small amount of **lead(II) nitrate** to one side of the dish. At the same time, have your partner add about the same amount of **potassium chloride** to the other side of the Petri dish.
3. Observe. Be patient! Write down what you see.

4

WRITING NET IONIC EQUATIONS



Lead(II) nitrate + potassium chloride →

1. Balanced Formula Equation (Grade 10 style!)
2. Complete ionic equation (show soluble salts as aqueous ions)
3. Net ionic equation (eliminate the spectator ions)